

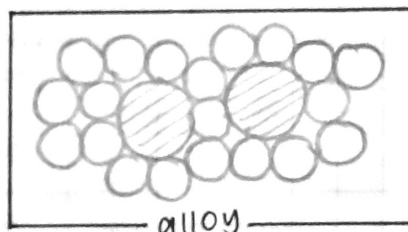
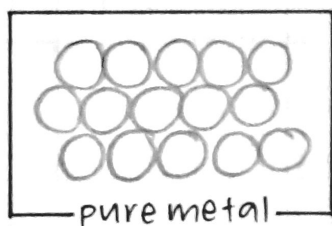
METALS

- physical properties

- high densities (except Group 1 metals)
- high melting and boiling points (except mercury)
- good conductors of heat and electricity
- shiny, ductile and malleable

ALLOYS

- used instead of pure metals which are soft and may corrode easily
- mixture of a metal and one or more elements, which may be metal/non metal
- harder because arrangement of atoms is disrupted by presence of atoms of different sizes, which prevents layers of atoms from sliding over each other easily



REACTIVITY SERIES

most reactive ↑	potassium	_____	please
	sodium	_____	stop
	calcium	_____	calling
	magnesium	_____	monkey
	aluminium	_____	and
	zinc	_____	zebra
	iron	_____	idiots
	lead	_____	leave
	copper	_____	come
	silver	_____	S
least reactive ↓	gold	_____	G

DISPLACEMENT REACTION

- takes place when a more reactive metal is placed in the salt solution of a less reactive metal
- more reactive metal displaces less reactive metal from its salt
- no reaction occurs when a less reactive metal is placed in the salt solution of a more reactive metal